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**CHEMISTRY**  
**SCIENCE Paper – 2**

(Two hours)

*Answers to this Paper must be written on the paper provided separately.*

*You will **not** be allowed to write during the first 15 minutes.*

*This time is to be spent in reading the Question Paper.*

*The time given at the head of this paper is the time allowed for writing the answers.*

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**Section I is compulsory. Attempt any four questions from Section II.**

*The intended marks for questions or parts of questions are given in brackets [ ].*

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**SECTION I (40 Marks)**

*Attempt **all** questions from this Section*

**Question 1**

- (a) Choose the correct answer from the options given below: [5]
- (i) An *electrolyte* which completely dissociates into ions is:
- A. Alcohol
  - B. Carbonic acid
  - C. Sucrose
  - D. Sodium hydroxide
- (ii) The most *electronegative element* from the following elements is:
- A. Magnesium
  - B. Chlorine
  - C. Aluminium
  - D. Sulphur

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**This Paper consists of 9 printed pages and 1 blank page.**

(iii) The reason for using *Aluminium* in the alloy duralumin is:

- A. Aluminium is brittle.
- B. Aluminium gives strength.
- C. Aluminium brings lightness.
- D. Aluminium lowers melting point.

(iv) The *drying agent* used to *dry HCl* gas is:

- A. Conc.  $\text{H}_2\text{SO}_4$
- B.  $\text{ZnO}$
- C.  $\text{Al}_2\text{O}_3$
- D.  $\text{CaO}$

(v) A hydrocarbon which is a *greenhouse gas* is:

- A. Acetylene
- B. Ethylene
- C. Ethane
- D. Methane

(b) Fill in the blanks with the choices given in brackets:

[5]

- (i) Conversion of *ethanol to ethene* by the action of *concentrated sulphuric acid* is an example of \_\_\_\_\_. (dehydration / dehydrogenation / dehydrohalogenation)
- (ii) When *sodium chloride* is heated with *concentrated sulphuric acid below  $200^\circ\text{C}$* , one of the products formed is \_\_\_\_\_. (sodium hydrogen sulphate / sodium sulphate / chlorine)
- (iii) *Ammonia* reacts with *excess chlorine* to form \_\_\_\_\_. (nitrogen / nitrogen trichloride / ammonium chloride)
- (iv) *Substitution reactions* are characteristic reactions of \_\_\_\_\_. (alkynes / alkenes / alkanes)
- (v) In Period 3, the *most metallic* element is \_\_\_\_\_. (sodium / magnesium / aluminium)

(c) Write a balanced chemical equation for each of the following reactions: [5]

- (i) Reduction of copper (II) oxide by hydrogen.
- (ii) Action of dilute sulphuric acid on sodium hydroxide.
- (iii) Action of dilute sulphuric acid on zinc sulphide.
- (iv) Ammonium hydroxide is added to ferrous sulphate solution.
- (v) Chlorine gas is reacted with ethene.

(d) State one observation for each of the following: [5]

- (i) Concentrated nitric acid is reacted with sulphur.
- (ii) Ammonia gas is passed over heated copper (II) oxide.
- (iii) Copper sulphate solution is electrolysed using copper electrodes.
- (iv) A small piece of zinc is added to dilute hydrochloric acid.
- (v) Lead nitrate is heated strongly in a test tube.

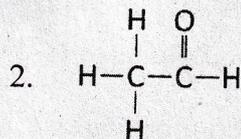
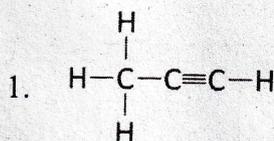
(e) (i) Calculate: [5]

- 1. The number of moles in 12g of oxygen gas. [O = 16]
- 2. The weight of  $10^{22}$  atoms of carbon.

[C = 12, Avogadro's No. =  $6 \times 10^{23}$ ]

(ii) Molecular formula of a compound is  $C_6H_{18}O_3$ . Find its empirical formula.

(f) (i) Give the IUPAC name of the following organic compounds: [5]



- (ii) What is the special feature of the structure of ethyne?
- (iii) Name the saturated hydrocarbon containing two carbon atoms.
- (iv) Give the structural formula of Acetic acid.
- (g) Give the appropriate term defined by the statements given below: [5]
- (i) The formula that represents the simplest ratio of the various elements present in one molecule of the compound.
- (ii) The substance that releases hydronium ion as the only positive ion when dissolved in water.
- (iii) The tendency of an atom to attract electrons towards itself when combined in a covalent compound.
- (iv) The process by which certain ores, specially carbonates, are converted to oxides in the absence of air.
- (v) The covalent bond in which the electrons are shared equally between the combining atoms.
- (h) Arrange the following according to the instructions given in brackets: [5]
- (i) K, Pb, Ca, Zn. (In the increasing order of the reactivity)
- (ii)  $Mg^{2+}$ ,  $Cu^{2+}$ ,  $Na^{1+}$ ,  $H^{1+}$  (In the order of preferential discharge at the cathode)
- (iii) Li, K, Na, H (In the decreasing order of their ionization potential)
- (iv) F, B, N, O (In the increasing order of electron affinity)
- (v) Ethane, methane, ethene, ethyne. (In the increasing order of the molecular weight) [H = 1, C = 12]



### Question 3

- (a) Name the particles present in: [3]
- (i) Strong electrolyte
  - (ii) Non- electrolyte
  - (iii) Weak electrolyte
- (b) Distinguish between the following pairs of compounds using the reagent given in the bracket. [3]
- (i) Manganese dioxide and copper (II) oxide. (using concentrated HCl)
  - (ii) Ferrous sulphate solution and ferric sulphate solution. (using sodium hydroxide solution)
  - (iii) Dilute hydrochloric acid and dilute sulphuric acid. (using lead nitrate solution)
- (c) Choose the method of preparation of the following salts, from the methods given in the list: [4]
- [List: A. Neutralization      B. Precipitation  
C. Direct combination      D. Substitution]
- (i) Lead chloride
  - (ii) Iron (II) sulphate
  - (iii) Sodium nitrate
  - (iv) Iron (III) chloride

### Question 4

- (a) Complete the following equations: [3]
- (i)  $S + \text{conc. HNO}_3 \rightarrow$
  - (ii)  $C + \text{conc. H}_2\text{SO}_4 \rightarrow$
  - (iii)  $\text{Cu} + \text{dil. HNO}_3 \rightarrow$

- (b) Write a balanced chemical equation for the preparation of: [3]
- (i) Ethene from bromoethane
  - (ii) Ethyne using calcium carbide
  - (iii) Methane from sodium acetate.
- (c) Name the following organic compounds: [4]
- (i) The compound with 3 carbon atoms whose functional group is a carboxyl.
  - (ii) The first homologue whose general formula is  $C_nH_{2n}$ .
  - (iii) The compound that reacts with acetic acid to form ethyl ethanoate.
  - (iv) The compound formed by complete chlorination of ethyne.

### Question 5

- (a) Give the chemical formula of: [3]
- (i) Bauxite
  - (ii) Cryolite
  - (iii) Sodium aluminate
- (b) Answer the following questions based on the extraction of aluminium from alumina by **Hall-Heroult's Process**.: [3]
- (i) What is the function of cryolite used along with alumina as the electrolyte?
  - (ii) Why is powdered coke sprinkled on top of the electrolyte?
  - (iii) Name the electrode, from which aluminium is collected.

(c) Match the alloys given in column I to the uses given in column II:

[4]

COLUMN I	COLUMN II
(i) Duralumin	A. Electrical fuse
(ii) Solder	B. Surgical instruments
(iii) Brass	C. Aircraft body
(iv) Stainless Steel	D. Decorative articles

### Question 6

(a) Identify the substances underlined:

[3]

- (i) The catalyst used to oxidise ammonia.
- (ii) The organic compound which when solidified, forms an ice like mass.
- (iii) The dilute acid which is an oxidizing agent.

(b) Copper sulphate solution reacts with sodium hydroxide solution to form a precipitate of copper hydroxide according to the equation:

[3]



(i) What mass of copper hydroxide is precipitated by using 200 gm of sodium hydroxide?

$$[\text{H} = 1, \text{O} = 16, \text{Na} = 23, \text{S} = 32, \text{Cu} = 64]$$

(ii) What is the colour of the precipitate formed?

(c) Find the **empirical formula** and the **molecular formula** of an organic compound from the data given below:

[4]

$$\text{C} = 75.92\%, \text{H} = 6.32\% \text{ and } \text{N} = 17.76\%$$

The vapour density of the compound is 39.5.

$$[\text{C} = 12, \text{H} = 1, \text{N} = 14]$$

### Question 7

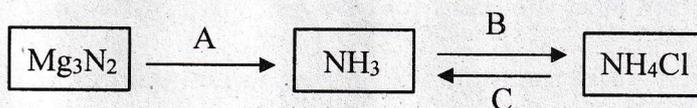
(a) Name the gas evolved in each of the following cases: [3]

(i) Alumina undergoes electrolytic reduction.

(ii) Ethene undergoes hydrogenation reaction.

(iii) Ammonia reacts with heated copper oxide.

(b) Study the flow chart given and give balanced equations to represent the reactions **A**, **B** and **C**: [3]



(c) Copy and complete the following table which refers to the **industrial method** for the preparation of ammonia and sulphuric acid: [4]

Name of the compound	Name of the process	Catalytic equation (with the catalyst)
Ammonia	(i) _____	(ii) _____
Sulphuric acid	(iii) _____	(iv) _____